

Chemical equations

Reactants: the substances you have at the beginning.

Products: the new substances you have when the reaction is complete.

Reactants → *Products*

or

Reactant + Reactant → *Product + Product*

Balancing Chemical Equations

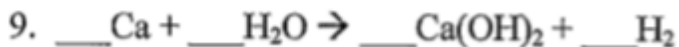
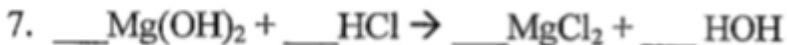
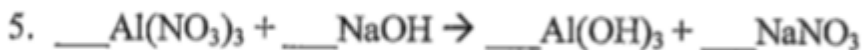
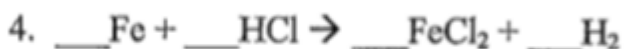
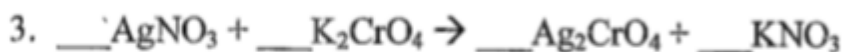
A chemical equation must show the same number of each type of atom on both sides of the equation.

After watching the video and reading pages 60-61 try your best to balance the following equations.

Ex: Original eqn.: $\text{Mg} + \text{O}_2 \rightarrow \text{MgO}$ Balanced: $2\text{Mg} + \text{O}_2 \rightarrow 2\text{MgO}$

There will be spots where there is a coefficient of 1, you can put a 1 or leave it blank.

Using only coefficients, fill in the blanks and balance the following equations.



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