

Counting atoms:

subscript: tells you the number of atoms in a molecule or the ratio of elements in a compound.

Name each of the elements in the following. Also tell how many atoms of each element are in the molecules or group of atoms.

1. F \_\_\_\_\_

2. N<sub>2</sub> \_\_\_\_\_

3. CH<sub>4</sub> \_\_\_\_\_

\_\_\_\_\_

4. CO \_\_\_\_\_

\_\_\_\_\_

5. Li<sub>2</sub>O \_\_\_\_\_

\_\_\_\_\_

6. H<sub>2</sub>CO<sub>3</sub> \_\_\_\_\_

\_\_\_\_\_

\_\_\_\_\_

Chemical equations

Reactants: the substances you have at the beginning.

Products: the new substances you have when the reaction is complete.

Reactants → Products

Reactant + Reactant → Product + Product

Coefficient: A number placed in front of a chemical formula in an equation. Ex: 2H<sub>2</sub>O

\*It tells you how many atoms or molecules of a reactant or product take part in the reaction.\*

1. 3OH \_\_\_\_\_

\_\_\_\_\_

2. 4CO<sub>2</sub> \_\_\_\_\_

\_\_\_\_\_

3. 3Al(NO<sub>3</sub>)<sub>3</sub> \_\_\_\_\_

\_\_\_\_\_